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POGIL Chemistry Teachers  
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Chemistry Answers

Calculating Ph Pogil Answer

Key  $\text{pH} = 14.000 - \text{pOH} =$

$14.000 - 2.187 = 11.813$  4)

-A solution is created by  
measuring  $3.60 \times 10^3$  moles  
of NaOH and  $5.95 \times 10^{-4}$   
moles of HCl into a

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container and then water is added until the final volume is 1.00 L. Calculating pH and pOH worksheet

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A Calculating pH - Manasquan  
Public Schools pOH =  
 $-\log[\text{OH}^-] = -\log(6.2 \times 10^{-5})$   
 $= 4.21$  pH = 14.00 - pOH =  
 $14.00 - 4.21 = 9.79$  6) +A  
-solution with a H  
concentration of  $1.00 \times 10^7$   
M is said to be neutral.  
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ph the sum of ph and poh is  
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for calculating ph the sum  
of ph and poh is always 14  
this is because the product



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of proton concentration and hydroxide concentration must always equal the equilibrium constant for the ionization of water which is equal to calculating ph

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Calculating Ph Packet

Answers Pogil

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 $= -\log(6.2 \times 10^{-5}) = 4.21$  pH  
 $= 14.00 - \text{pOH} = 14.00 - 4.21$   
 $= 9.79$  6) +A - solution with  
a H concentration of  $1.00 \times$   
 $10^{-7}$  M is said to be neutral.

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