

Chemistry 2nd Semester Study Guide

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1) calculate moles of each element (if given mass) 2) divide each mole quantity by the smallest mole value from step 1. 3) use whole number values from step 2 as the subscripts to the original equation to write the empirical formula. Percent Composition. -the relative amounts of each element in a compound.

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Chemistry - 2nd Semester Practice Test Thermochemistry (Ch.11): Multiple Choice 24) What is the amount of heat required to raise the temperature of 200.0 g of aluminum by 10\u00b0C? cal (specific heat of aluminum = 0.21) g x \u00b0C 24) A) 4200 cal B) 420 cal C) 42 000 cal D) 420 000 cal E) none of the above

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2nd Semester Chemistry-575 Final Exam Review Multiple Choice Identify the choice that best completes the statement or answers the question. ____ 1. How many moles of tungsten atoms are in 4.8 \u00d7 1025 atoms of tungsten? a. 8.0 \u00d7 102 moles c. 1.3 \u00d7 10-1 moles b. 8.0 \u00d7 101 moles d. 1.3 \u00d7 10-2 moles ____ 2.

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Chapter 2 History of chemistry before 1701 Chemistry had a bad reputation because of alchemy, which was a pseudoscience that claimed to convert cheap metals into gold. Alchemy led to the discovery of many elements. 1500 - George Bauer discovered metallurgy, extracting pure metal out of a mix 1600s - Robert Boyle studied the behavior of gases and is responsible for the theory of elements ...