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Laboratory manual for physical chemistry where T0 is the freezing point of the solvent. If BC in Fig. 1st-1 is taken to be a straight line, the slope dpA/dT may by replaced by CB/BD, i.e., by (pA - ps)/ ΔT, so that pp T Hp RT AsmvA 0 2 – = δ Δ 0 Eq. 1st.4 The slope of the vapor pressure curve of the solid solvent is also given by another form of the

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The dilution formula allows you to determine the molarity of a solution after it has been diluted. The formula is: M 1V 1=M 2V 2 Where M 1 and V 1 are the molarity and volume before a dilution, and M 2 and V 2 are the molarity and volume after a dilution. Consider this example from the 2006 NECTA: "Hydrochloric acid solution made by diluting

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Laboratory Manual of Practical Inorganic II Chemistry. Figure 1.1 Packing models of (a) side-by-side layer, (b) hexagon layer, (c) simple cubic layer, (d) body centered cubic. Determine the maximum amount of balls that can touch another one ball (center- ball) in the same layer for each arrangement.

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