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Chemistry RHS – Mr. Moss 1. Define the following: a. Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction. b. Mole –The SI unit used to measure the amount of a substance that contains 6.02×10^{23} atoms of that substance.

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Chemistry RHS Mr. Moss

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- Volume 11.2 Stoichiometric Calculations Concepts Study Guide for Chapter 11 Stoichiometry 368
- Chapter 11 • Stoichiometry Section 11.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation.

Stoichiometry Chapter 11 Study Guide Answer Key

TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 – Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio

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Stoichiometric Calculations In your textbook, read about mole-to-mole conversion. Read the following passage and then solve the problems. In the equation that follows each problem, write in the space provided the mole ratio that can be used to solve the problem.

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